

The Structures of Vanadium Oxide Cluster–Ethene Complexes. A Combined IR Multiple Photon Dissociation Spectroscopy and DFT Calculation Study

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Vanadium is one of the most important transition metals in supported metal oxide catalysis, where it is mainly used as an oxidation catalyst.¹ However, in most of these catalytic reactions, the mechanisms on a molecular scale and the relevant active sites are still under discussion. Gas-phase vanadium oxide clusters have been suggested as model systems for active sites, and reactions with small organic molecules have been studied experimentally using mass spectrometric techniques.^{2–4} Evidence for oxygen transfer reactions has been found.^{2,3} On the basis of the structure–reactivity relationship obtained from DFT calculations, a radical center mechanism has been proposed to be responsible for the size-selective oxygen transfer.³ The complexity and importance of those reactions cause a continued interest in the mechanisms of these reactions as well as in the structures of the oxide cluster–hydrocarbon complexes.

Most of the knowledge about the structure–reactivity relationships for gas-phase clusters has been gained indirectly from experimental chemical reactivity, in some cases combined with quantum chemical calculations. More direct structural information can be obtained from vibrational spectroscopy. IR multiple photon dissociation (IR-MPD) with line tunable CO₂ lasers has been used to obtain IR spectral information for some metal cluster complexes in a limited spectral range around 10 μm.⁵ Using a free electron laser (FEL) as an intense and widely tunable IR source, IR-MPD spectra have been measured for oxide clusters of vanadium, niobium, and tantalum,^{6–8} as well as for some complexes of metal clusters.^{9,10} For some small oxide clusters, this led to structural identification through comparison with DFT calculations.

Here, we report on IR-MPD experiments that are performed in a molecular beam setup on cluster ions directly emitted by a pulsed laser vaporization cluster source that is combined with a reaction channel. An additional pulsed valve delivers the ethene to the reaction channel, where the clusters react near thermal energy. The setup is described in more detail elsewhere.^{7,10,11} The clusters and their complexes form a molecular beam that is overlapped collinearly with the intense infrared beam of the Free Electron Laser for Infrared eXperiments (FELIX).^{11,12} When the laser frequency is resonant with an IR-active vibration of a cluster, the cluster can absorb several photons and subsequently undergo fragmentation. The changes in the cationic cluster mass distribution are monitored as a function of FELIX wavelength. The mechanism of IR multiple photon absorption has been discussed recently in the context of IR resonance enhanced multiple photon ionization (IR-REMPI) spectroscopy of clusters¹¹ as well as IR-MPD of gas-phase molecular ions.¹³

The use of IR spectroscopy for structural identification depends on the ability of theoretical methods to accurately predict structural and dynamical properties as well as energetics. Therefore, the structural assignment is a joint experimental and theoretical effort. Recent work on structural properties of cationic vanadium oxide clusters and their reactivity with ethene has shown that the DFT methods, in particular, the hybrid gradient corrected functionals (B3LYP) in combination with sufficiently large basis sets of triple-ζ quality (TZVP), allow accurate determination of the structural features and also of the structure–reactivity relationship.³ Structural properties of anionic and neutral¹⁴ as well as cationic¹⁵ V_xO_y have been calculated using the same procedure. These findings gave us confidence to extend this approach to the interpretation of vibrational IR spectra of cationic vanadium oxide–ethene complexes reported in this work. For this purpose, the vibrational frequencies and IR intensities have been calculated within the harmonic approximation and are compared here to experimental data.

IR-MPD spectra for the complexes of V₂O₅⁺ and V₂O₆⁺ with C₂H₄ are shown in Figure 1 together with the calculated spectra of some isomers. Both oxide cluster cations are open-shell systems; in V₂O₅⁺, the unpaired electron is located at a terminal V–O• group, whereas V₂O₆⁺ contains a superoxo group (formally O₂^{•-}) that can be identified via its characteristic IR absorption at ~1150 cm⁻¹. Good agreement is achieved between the experimental IR spectra of V₂O₅C₂H₄⁺ (**I**) and V₂O₆C₂H₄⁺ (**II**) and the calculated spectra of the isomers **Ic** and **IIa**, respectively (**Ia** is incompatible due to the observed fragmentation behavior; see below). In principle, the agreement could be improved if one would introduce an optimal scaling factor for all modes involving V–O stretches. Because of the lack of reference data for similar gas-phase molecules, we used here a rather arbitrary factor of 0.95 for all vibrational frequencies. While **IIa** is the lowest energy isomer of V₂O₆C₂H₄⁺, **Ic** is 1.98 eV higher in energy than the minimum energy isomer **Ia** obtained for V₂O₅C₂H₄⁺.

In addition to the IR spectral information, fragmentation pathways can be identified by analyzing the frequency dependence of the intensity increase of the fragment ions. For complex **I**, the observed fragmentation channel is [C₂H₄] loss, and for **II** loss of [C₂H₄O₂] (presumably successive loss of C₂H₄ and O₂) is observed. The calculated energetics of the V₂O₅⁺/C₂H₄ system are shown in Figure 2. A reaction between C₂H₄ and V₂O₅⁺ could follow the lowest energy path, but then the high intrinsic energy should cause immediate formation of the acetaldehyde as oxidation product. If **Ia** would nonetheless form for some reason, pumping it with IR light is expected to induce rearrangement to an acetaldehyde complex of V₂O₄⁺ (**Id**), combined with loss of acetaldehyde. Neither the strong ν_{CO} absorption of **Id** (at 1584 cm⁻¹) nor loss of [C₂H₄O] has been experimentally observed. Although the IR spectra of **Ia**

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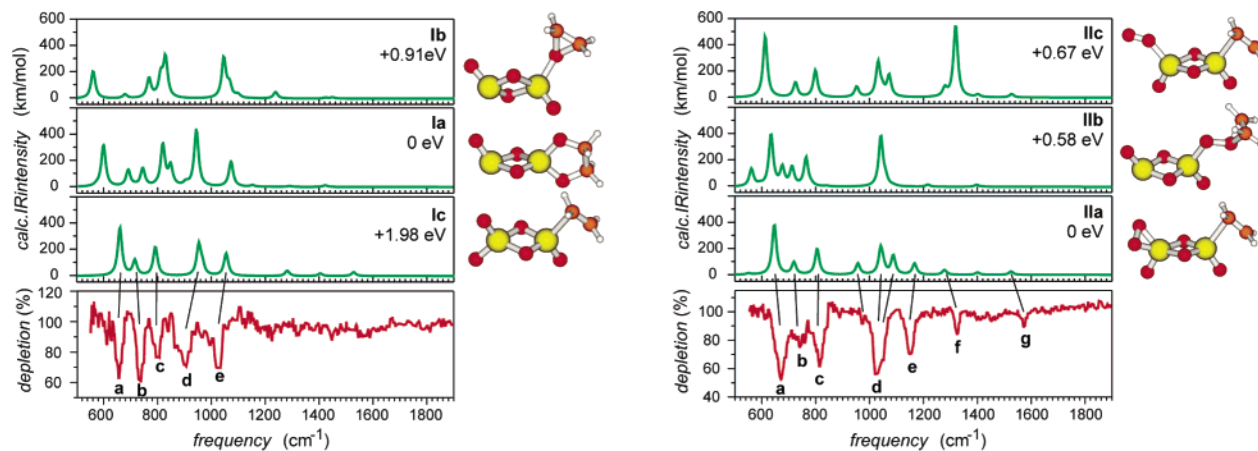


Figure 1. Infrared multiple photon dissociation (IR-MPD) spectra of $V_2O_5C_2H_4^+$ (**I**) and $V_2O_6C_2H_4^+$ (**II**) together with calculated spectra for several isomers.

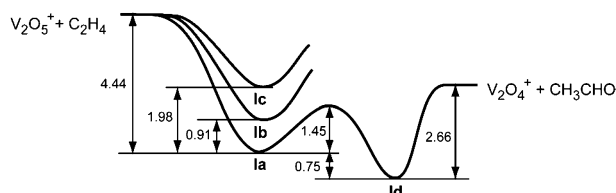


Figure 2. Schematic diagram showing the relative energies (in eV) of four different isomers of $V_2O_5C_2H_4^+$ (**Ia–d**) and the energetics of the ethene oxidation on **Ia** leading to acetaldehyde (compare ref 3).

and **Ic** are similar, the observed loss of $[C_2H_4]$ thus makes the presence of **Ic** more probable, and, instead of forming the energetically more stable O–C bonds, the ethene subunit seems to be bound directly to the vanadium atom with the lowest O–coordination in both **Ic** and **IIa**.

From these findings, the question arises of why the vanadium oxide clusters behave differently during reaction with ethene in our experiment than at the conditions of mass selected cluster reactivity measurement. There,³ theory predicts the bonding of ethene to an O-site as the initial step for the experimentally observed reduction of $V_2O_5^+$. However, we find the ethene to bind to a V-site. We suggest that this is caused by the influence of an electrostatic effect on the geometry of the ion molecule reaction. The positive charge is located in the oxide cluster at the lowest coordinated V-atom and can orient the ethene molecule. In that way, an intermediate ion–molecule complex is formed which is not necessarily the lowest energy structure. Thus, the here identified structures resemble the primary collision complexes that are formed at thermal collision energies. In a mass selected cluster reactivity measurement, the collision energies are typically at least a few electronvolts. Even if this energy is only partly transferred into internal energy, barriers can be overcome, and more stable products can be formed.

The comparison of the experimental and calculated spectra as well as (qualitative) investigations of the laser power dependence of the dissociation spectra gave us no evidence for the presence of different structural isomers of **I** or **II**. For **II**, the calculated IR spectra of the higher energetic isomers differ more from the experimental spectrum than for the most stable isomer, thus supporting the assignment to **IIa**. For **I**, the observed loss of $[C_2H_4]$ complements the ambiguous spectroscopic observations, giving evidence for the presence of isomer **Ic**. In both complexes, the ethylene molecule binds to the initially positively charged V atom.

In conclusion, we have presented infrared spectra of complexes

of transition metal oxide clusters with a hydrocarbon molecule. We conclude that the cationic nature of the vanadium oxide clusters may have a dominating influence over the reactions under the applied conditions.

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Supporting Information Available: Structures, vibrational frequencies, and IR intensities for **Ia–d** and **IIa–c**, and band assignments for **I** and **II** (PDF). This material is available free of charge via the Internet at <http://pubs.acs.org>.

References

- (1) Weckhuysen, B. M.; Keller, D. E. *Catal. Today* **2003**, *78*, 25.
- (2) Zemski, K. A.; Justes, D. R.; Castleman, A. W., Jr. *J. Phys. Chem. B* **2002**, *106*, 6136. Bell, R. C.; Castleman, A. W., Jr. *J. Phys. Chem. A* **2002**, *106*, 9893.
- (3) Justes, D. R.; Mitrić, R.; Moore, N. A.; Bonačić-Koutecký, V.; Castleman, A. W., Jr. *J. Am. Chem. Soc.* **2003**, *125*, 6289.
- (4) Fielicke, A.; Rademann, K. *Phys. Chem. Chem. Phys.* **2002**, *4*, 2621. Dinca, A.; Davis, T. P.; Fisher, K. J.; Smith, D. R.; Willett, G. D. *Int. J. Mass Spectrom.* **1999**, *182/183*, 73.
- (5) Zakin, M. R.; Brickman, R. O.; Cox, D. M.; Reichmann, K. C.; Trevor, D. J.; Kaldor, A. *J. Chem. Phys.* **1986**, *85*, 1198. Knickelbein, M. B. *Annu. Rev. Phys. Chem.* **1999**, *50*, 79. Koretsky, G. M.; Knickelbein, M. B.; Rousseau, R.; Marx, D. *J. Phys. Chem. A* **2001**, *105*, 11197. Dietrich, G.; Krückeberg, S.; Lützenkirchen, K.; Schweikhard, L.; Walther, C. *J. Chem. Phys.* **2000**, *112*, 752.
- (6) Asmis, K. R.; Brümmer, M.; Kaposta, C.; Santambrogio, G.; von Helden, G.; Meijer, G.; Rademann, K.; Wöste, L. *Phys. Chem. Chem. Phys.* **2002**, *4*, 1101. Fielicke, A.; Meijer, G.; von Helden, G. *Eur. Phys. J. D* **2003**, *24*, 69.
- (7) Fielicke, A.; Meijer, G.; von Helden, G. *J. Am. Chem. Soc.* **2003**, *125*, 3659.
- (8) Asmis, K. R. 2002, private communication. IR spectra of $V_2O_x^+$ ($x = 2–6$) were reported at ISSPIC 11 in Strasbourg, France, September 2002.
- (9) Simard, B.; Dénoimée, S.; Rayner, D. M.; van Heijnsbergen, D.; Meijer, G.; von Helden, G. *Chem. Phys. Lett.* **2002**, *357*, 195.
- (10) Fielicke, A.; von Helden, G.; Meijer, G.; Simard, B.; Dénoimée, S.; Rayner, D. M. *J. Am. Chem. Soc.* **2003**, *125*, 11184.
- (11) von Helden, G.; van Heijnsbergen, D.; Meijer, G. *J. Phys. Chem. A* **2003**, *107*, 1671.
- (12) Oepts, D.; van der Meer, A. F. G.; van Amersfoort, P. W. *Infrared Phys. Technol.* **1995**, *36*, 297.
- (13) Oomens, J.; Tielens, A. G. G. M.; Sartakov, B.; von Helden, G.; Meijer, G. *Astrophys. J.* **2003**, *591*, 968.
- (14) Vyboishchikov, S. F.; Sauer, J. *J. Phys. Chem. A* **2000**, *104*, 10913. Vyboishchikov, S. F.; Sauer, J. *J. Phys. Chem. A* **2001**, *105*, 8588.
- (15) Sauer, J.; Vyboishchikov, S. F. Poster presentation, Humboldt-Universität zu Berlin, Germany, February 2002.

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